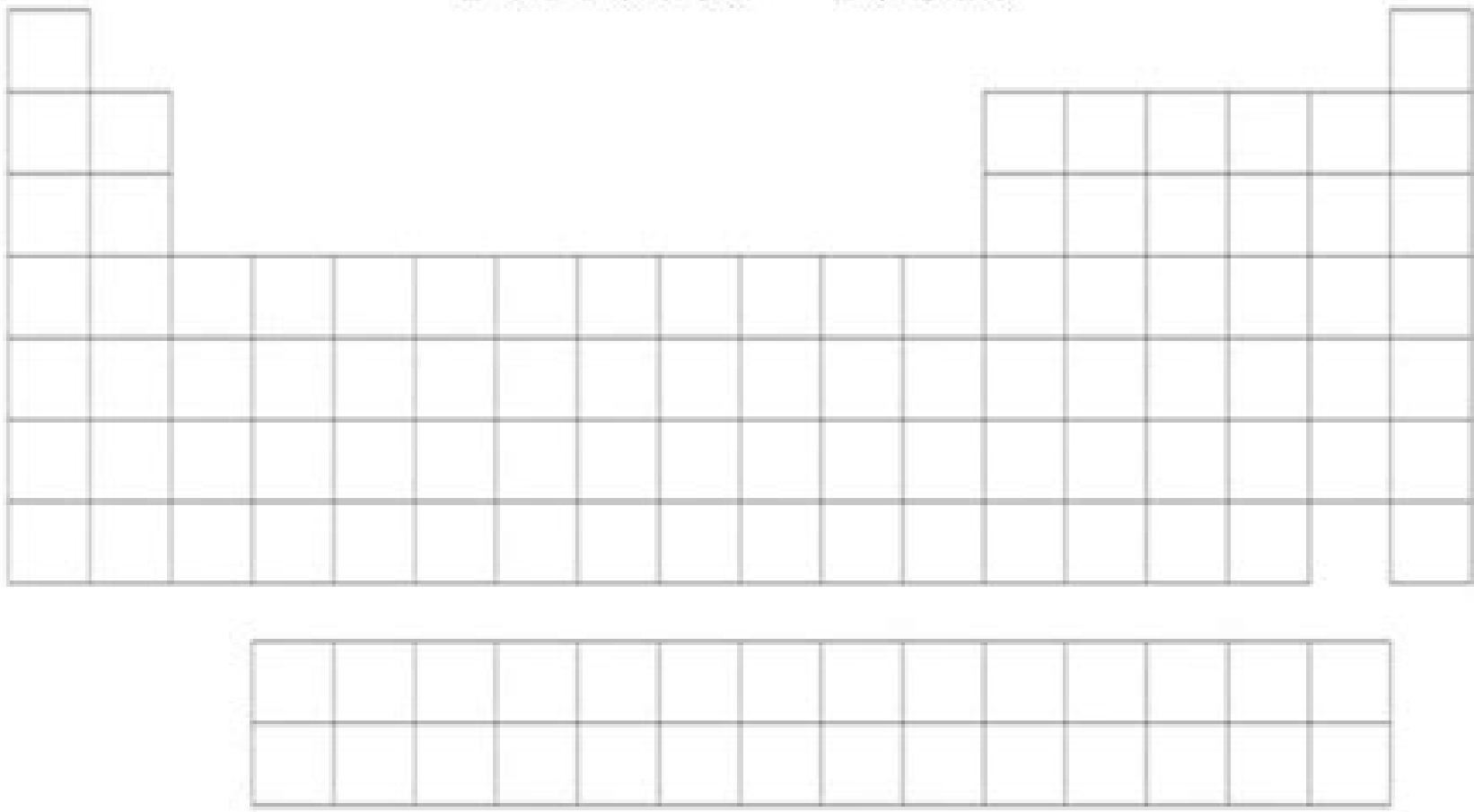


I'm not robot!

Periodic Table



Kenzie Mayhew © 2011

Name_____

Scientific Notation

Scientists very often deal with very small and very large numbers, which can lead to a lot of confusion when counting zeros. We can express these numbers as powers of 10.

Scientific notation takes the form of $M \times 10^n$ where $1 \leq M < 10$ and n represents the number of decimal places to be moved. Positive n indicates the standard form is a large number. Negative n indicates a number between zero and one.

Example 1: Convert 1,500,000 to scientific notation.
Move the decimal point so that there is only one digit to its left, for a total of 6 places.
 $1,500,000 = 1.5 \times 10^6$

Example 2: Convert 0.000025 to scientific notation.
For this, move the decimal point 5 places to the right.
 $0.000025 = 2.5 \times 10^{-5}$

(Note that when a number starts out less than one, the exponent is always negative.)

Convert each number to scientific notation.

1. 0.005 = _____

2. 5.050 = _____

3. 0.0008 = _____

4. 1,000 = _____

5. 1,000,000 = _____
6. 0.25 = _____

7. 0.025 = _____

8. 0.0025 = _____

9. 500 = _____

10. 5,000 = _____

Convert each number to standard notation.

11. $1.5 \times 10^3 =$ _____

12. $1.5 \times 10^{-3} =$ _____

13. $3.75 \times 10^{-2} =$ _____

14. $3.75 \times 10^2 =$ _____

15. $2.2 \times 10^5 =$ _____
16. $3.35 \times 10^{-1} =$ _____

17. $1.2 \times 10^{-4} =$ _____

18. $1 \times 10^4 =$ _____

19. $1 \times 10^{-1} =$ _____

20. $4 \times 10^3 =$ _____

Quarter 4 - Matter and its Interaction

MODULE 1 - BEHAVIOR OF GASES

Activity 3 - Charles' Law

Procedure:

1. Prepare 3 set-ups using beakers/basin.

a. Set-up A (hot water)

b. Set-up B (tap water)

c. Set-up C (cold water)

2. Inflate a balloon (it must be spherical).

3. After inflating, draw a line around the widest (or middle portion) of the balloon to be consistent with where to measure the circumference.

(Note: The same balloon will be used throughout the experiment. Thus, do not overinflate the balloon.)

4. Predict first what will happen to the size of the balloon before submerging the balloons into the different set-ups. Write your predictions on Table 1.

5. Measure the circumference of the balloon.

6. Get the temperature of the water in the basin (Set up A).

7. Place inflated balloon in set-up A for 2-3 minutes.

8. Measure the circumference of the balloon once more.

9. Perform two more trials and get the average of the results.

10. Repeat step 5-9 using Set B and C.

11. Record your observations on Table 2.

Group no. _____

Group name: _____

Group members:

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

Set up / trials	Temp (°C)	Average Temp (°C)	Average Temp (K)	Circumference of the balloon (cm)			Volume of the balloon after placing in a set up.	Average volume
				Before	after	diff		
Warm	1							
	2							
	3							
Tap water	1							
	2							
	3							
Cold water	1							
	2							
	3							

- Q1. What happens to the size of the balloon as the temperature decreases?
- Q2. How does the temperature relate to the volume of gas in the balloon?

Science Department - Foundation	Science Department - Foundation
Extended writing task: Acids, alkalis and indicators	Extended writing task: Alloys and metal properties Extended writing task: Acids and neutralisation
Book description	Book description
1. This book is about acids and alkalis.	1. This book is about alloys and metal properties.
2. It explains how acids and alkalis react.	2. It explains how alloys are made.
3. It shows how to test for acids and alkalis.	3. It shows how to test for different metals.
4. It gives examples of acids and alkalis.	4. It gives examples of different alloys.
5. It explains why acids and alkalis are important.	5. It explains why alloys are important.



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